Contents:
(3) H20
Enhancer Chips
per package.

Water Tastes Better!

H2O Enhancer Chip

Made in the U.S.A.

The H2O Enhancer Chip is a **non-drug** product that was developed as an enhancement device to improve Hydrogen bonding in water. This device is used on the exterior of any water container.

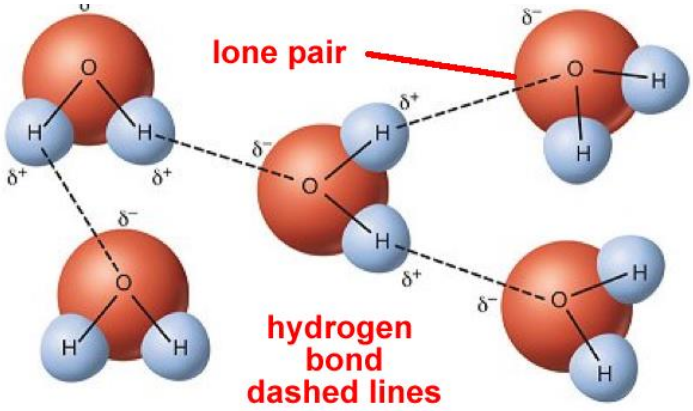
Drinking Water Helps Maintain the Balance of Body Fluids. Your body is composed of about 60% to 65% of water. The functions of these bodily fluids include digestion, absorption, circulation, creation of saliva, transportation of nutrients, and maintenance of body temperature.

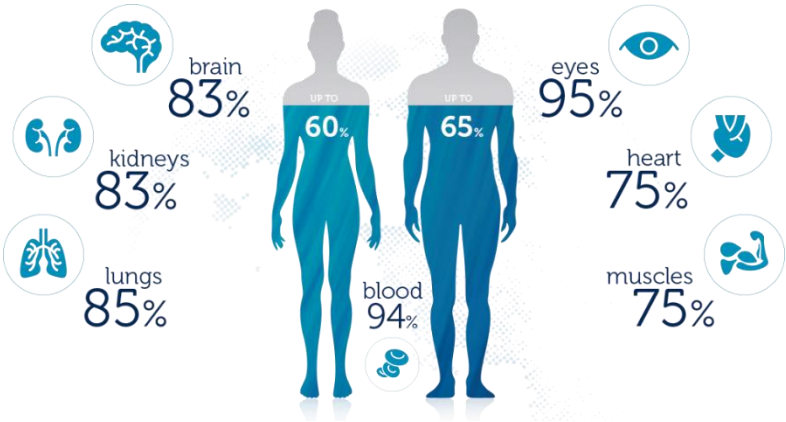
Water makes up more than two thirds of human body weight, and without water, we would die in a few days. The human brain is made up of 83% water, blood is 94% and lungs 85%. A mere 2% drop in our body's water supply can trigger signs of dehydration: fuzzy short-term memory, trouble with basic math, and difficulty focusing on smaller print, such as a computer screen.

Mild dehydration is also one of the most common causes of daytime fatigue. An estimated seventy-five percent of Americans have mild, chronic dehydration. Pretty scary statistic for a developed country where water is readily available through the tap or bottle water. The H2o Enhancer Chip is programmed to encourage water molecules to be absorbed better into the cells of your body

Hydrogen Bonding Example

These bonds are generally stronger than ordinary dipole-dipole and dispersion forces, but weaker than true covalent and ionic bonds.





Water Facts: The most important thing in the universe. Water is a base for all life. Scientists are searching water across the universe for possible life forms. Water takes about 71% of earth's surface. Around two-thirds of our body is made up of water, aids in nearly all processes of our body. It helps to maintain our body health in many ways. Each cell in our body has more than 50% of water in it. Water is the main constituent of life form.

H2O Enhancer Chip

Instructions: The H2O Enhancer Chip is used on any water container, up to (1) gallon in capacity. The H20 Enhancer Chip is electronically programmed to enhance the bonding structure to improve better hydration. It was developed to support the body's ability to hydrate more effectively. Made to last for 30 days, then discard. Apply another chip for additional 30 days to enhance hydration.

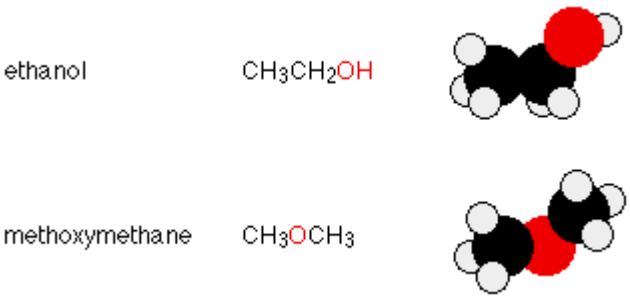
Water consumption directly affects energy levels. Just a 5% drop in body fluids can cause a 30% reduction in energy; the problem is not uncommon. By some estimates, two-thirds of the population suffers from some degree of chronic hydration. Rather than drink more water, many people turn to stimulants like caffeine and sugar to artificially boost their energy, quickly leading to a downward spiral. Caffeine and sugar are diuretics and cause your body to lose water, resulting in a further loss of energy and a dependency on artificial energy.

(Contents: 3 per package)

Hydrogen bonding in alcohols

An alcohol is an organic molecule containing an -OH group. Any molecule which has a hydrogen atom attached directly to an oxygen or a nitrogen is capable of hydrogen bonding. Such molecules will always have higher boiling points than similarly sized molecules which don't have an -O-H or an -N-H group. The hydrogen bonding makes the molecules "stickier", and more heat is necessary to separate them.

Ethanol, CH_3CH_2-O-HCH3CH2−O−H, and methoxymethane, CH_3-O-CH_3CH3−O−CH3, both have the same molecular formula, C_2H_6OC2H6O.



They have the same number of electrons, and a similar length to the molecule. The van der Waals attractions (both dispersion forces and dipole-dipole attractions) in each will be much the same. However, ethanol has a hydrogen atom attached directly to an oxygen - and that oxygen still has exactly the same two lone pairs as in a water molecule. Hydrogen bonding can occur between ethanol molecules, although not as effectively as in water. The hydrogen bonding is limited by the fact that there is only one hydrogen in each ethanol molecule with sufficient $\delta+$ charge.

In methoxymethane, the lone pairs on the oxygen are still there, but the hydrogens are not sufficiently $\delta+$ for hydrogen bonds to form. Except in some rather unusual cases, the hydrogen atom has to be attached directly to the very electronegative element for hydrogen bonding to occur. The boiling points of ethanol and methoxymethane show the dramatic effect that the hydrogen bonding has on the stickiness of the ethanol molecules:

ethanol (with hydrogen bonding)	78.5°C
methoxymethane (without hydrogen bonding)	-24.8°C

The hydrogen bonding in the ethanol has lifted its boiling point about 100°C. It is important to realize that hydrogen bonding exists in addition to van der Waals attractions. For example, all the following molecules contain the same number of electrons, and the first two are much the same length. The higher boiling point of the butan-1-ol is due to the additional hydrogen bonding.